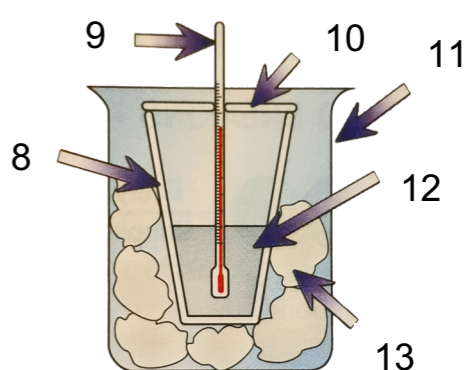
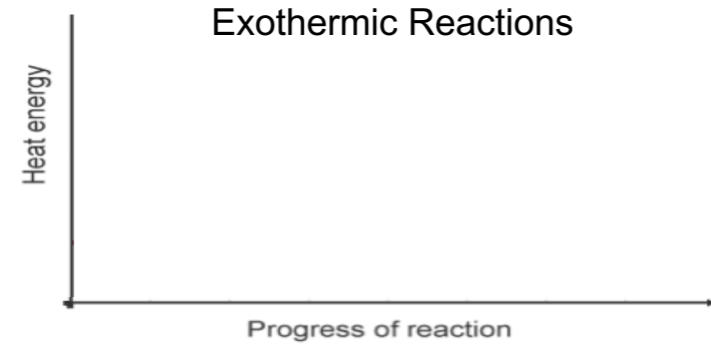


CHEMISTRY – YEAR 9 – Energy Changes Foundation

A ENERGY CHANGES OF REACTIONS		
1	Conservation of Energy	Energy is never destroyed It is transferred to surroundings or products.
2	Breaking Bonds	Energy is absorbed (endothermic)
3	Making Bonds	Energy is released (exothermic)
4	Overall Energy Change	= energy absorbed – energy released (LEFT-RIGHT)
5	Exothermic Reactions	Energy transferred to surroundings Surroundings heat up e.g. combustion
6	Endothermic Reactions	Energy transferred from surrounding Surroundings cool down e.g. sports injury pack
7	Activation Energy	Minimum energy required to start a reaction

A ENERGY CHANGES OF REACTIONS		
8	Polystyrene Cup	
9	Thermometer	
10	Lid	
11	Large Beaker	
12	Reaction Mixture	
13	Cotton Wool	
14	Explain Exothermic	Less energy required to break bonds in reactant than is released making bonds in product
15	Explain Endothermic	More energy required to break bonds in reactant than is released making bonds in product

B REACTION PROFILES	
1	<p>Exothermic Reactions</p> 
2	<p>Endothermic Reactions</p> 